



# TEST PAPER: CHEMISTRY

**Time: 45 Minutes**

**Class: C.B.S.E 10**

**Max. Marks: 30 marks**

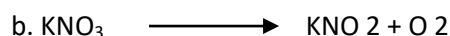
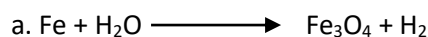
**Date: 2<sup>nd</sup> May 2018**

**Marking Scheme:** Three questions carry 10 marks each. Each question has 3 sub-parts. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.

## Question 1

A. Differentiate between combination and decomposition reactions with example

B. Balance the following chemical equations:



C. Explain exothermic and endothermic reactions with example.

## Question 2

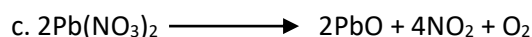
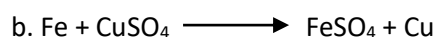
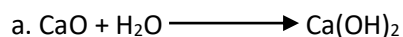
A. Translate following statements into chemical equations and then balance them:

a. Hydrogen gas combines with nitrogen to form ammonia

b. Hydrogen sulphide gas burns in air to give water and sulphur dioxide.

c. Aluminium reacts with hydrochloric acid to give aluminium chloride and hydrogen gas.

B. Identify the type of reaction for the following:



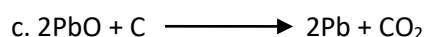
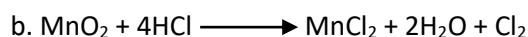
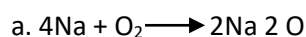
C. Define following terms:

a. Corrosion      b. Rancidity

## Question 3

A. Explain electrolysis of water with equations occurring at cathode and anode.

B. Identify the substances that are oxidised and that are reduced in the following reactions:



C. Substance X is used for white washing; it reacts with  $\text{CO}_2$  to form limestone (Y). Identify X and Y. Also, give balanced equation involved.

