



TEST PAPER: CHEMISTRY

Time: 45 Minutes

Class: I.C.S.E 9

Max. Marks: 30 marks

Date: 2nd May 2018

Marking Scheme: Three questions carry 10 marks each. Each question has 3 sub-parts. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.

Question 1

- A. Why metals lose electrons and non-metals gain electrons.
- B. What are ions? Represent ions if metal A loses two electrons and three electrons respectively.
- C. State Modern Periodic law. How many periods and columns are there in modern periodic table.

Question 2

- A. Give the names of group-II members with their characteristics.
- B. Give the electronic configuration of following elements:
 - a. Cs(55)
 - b. Sr(38)
 - c. Cl(17)
- C. Differentiate between alkali metals (group-I) and halogens (group-17).

Question 3

- A. State the members of group-18 with their configuration. Why the members are called inert gases?
- B. Define following terms:
 - a. Periodicity
 - b. Ionization energy
- C. Why group-I members are called alkali metals.