TEST PAPER: CHEMISTRY

Time: 50 Minutes **Class:** CBSE 10th

Max. Marks: 30 Marks **Date:** 13th June, 2018

Marking Scheme: *Three questions carry 10 marks each. Questions have 3 subparts each. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.*

Question 1:

A. Give reasons for the following:

a. Acids show acidic behavior in presence of water.
b. Oil and fat containing food items are flushed with nitrogen.
c. Iron articles and machinery are painted.

B. Define neutralization and peptisation reaction with examples.

C. Explain following with balanced chemical equation:

a. A metallic compound A reacts with dilute hydrochloric acid to produce effervescence. The gas evolved extinguishes a burning candle.
b. A metal A reacts with dilute hydrochloric acid to produce effervescence. The gas evolved burns with a pop sound.
c. A metallic oxide A reacts with water to form an alkali.
d. A non-metallic oxide A used to extinguish fire reacts with water to form an acid.

Question 2:

A. Are plants and animals pH sensitive. How pH change causes tooth decay.

B. Explain the nature of following salts obtained with the help of equations:

a. A salt of strong acid and strong base.
b. A salt of strong acid and weak base.
c. A salt of weak acid and strong base.

C. Give the balanced chemical equation for the following word equation:

a. Dilute hydrochloric acid reacts with zinc granules.
b. Dilute hydrochloric acid reacts with magnesium ribbon.
c. Dilute sulphuric acid reacts with iron fillings.
d. Dilute sulphuric acid reacts with aluminium powder.

Question 3:

A. State 3 between combination and decomposition reaction with examples.

B. Name the acids present in following substances:

Curd, ant sting, orange, tamarind, tomato and vinegar

C. Balance the following chemical equations:

a. $C + H_2SO_4 \longrightarrow CO_2 + H_2O + SO_2$
b. $C_2H_6 + O_2 \longrightarrow CO_2 + H_2O$
c. $Fe + H_2O \longrightarrow Fe_3O_4 + H_2$
d. $MnO_2 + HCl \longrightarrow MnCl_2 + H_2O + Cl_2$