



TEST PAPER: CHEMISTRY

Time: 50 Minutes

Class: CBSE 9th

Max. Marks: 30 Marks

Date: 13th June, 2018

Marking Scheme: Three questions carry 10 marks each. Questions have 3 subparts each. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.

Question 1:

A. Define the following terms:

- Crystallization*
- Tindall effect*
- Colloids*

B. Differentiate between homogenous and heterogenous mixture.

C. How can we separate a mixture of acetone and water? Explain with the help of diagram.

Question 2:

A. Solve the following:

- A solution contains 16 g of solute in 104 g of water. Calculate the concentration of solution in terms of mass by mass percentage of the solution.
- A solution contains 60 g of solute in 120 ml of solution. Calculate the concentration in terms of mass by volume of the solution.
- What is the mass by mass percentage of sodium hydroxide in a solution that is made by dissolving 8 g NaOH in 50 g of water?
- What is mass by volume percentage of the solution having 2.50 g of KCl dissolved in 50 ml of solution?

B. Give the dispersed phase and dispersion medium for the following colloids:

Fog, jelly, butter, milk and shaving cream

C. Explain the factors affecting evaporation. How does evaporation causes cooling?

Question 3:

A. Give two reasons to justify;

- Water at room temperature is a liquid.*
- An iron almira is a solid at room temperature.*

B. State 3 differences between evaporation and boiling.

C. Give reasons for the following:

- Water droplets on the outer surface of a glass containing ice cold water.*
- Water kept in earthen pot becomes cool during summer.*