



TEST PAPER: CHEMISTRY

Time: 50 Minutes

Class: ICSE 8th

Max. Marks: 30 Marks

Date: 13th June, 2018

Marking Scheme: Three questions carry 10 marks each. Questions have 3 subparts each. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.

Question 1:

A. Write the chemical formula of the following compounds:

- a. Potassium permanganate
- b. Magnesium nitride
- c. Aluminium sulphate
- d. Calcium phosphate
- e. Copper (I) sulphide
- f. Aluminium hydroxide

B. State the law of conservation of matter. Give a reason why an equation is balanced with reference to the law of conservation of matter.

C. Give the names of the following compounds:

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|------------------|-------------|
| a. Na_2ZnO_2 | g. Na_2O |
| b. $Mg(NO_3)_2$ | h. $ZnCl_2$ |
| c. $Ca(HCO_3)_2$ | |
| d. Al_2S_3 | |
| e. $Pb(OH)_2$ | |
| f. $K_2Cr_2O_7$ | |

Question 2:

A. Differentiate between liquefaction and solidification.

B. Define the following terms:

- a. Radicals
- b. Sublimation

C. Write the variable valencies of following metallic elements with their names:

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|-----------|---------|------------|
| a. Copper | b. Iron | c. Mercury |
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Question 3:

A. Give the valencies of following radicals with symbol:

Aluminate, sulphate, sulphite, bisulphate, bisulphite, nitride, nitrate

B. Balance the following Chemical equations:

- a. $C + H_2SO_4 \rightarrow CO_2 + H_2O + SO_2$
- b. $N_2 + H_2 \rightarrow NH_3$
- c. $FeS + HCl \rightarrow FeCl_2 + H_2S$

C. Write balanced chemical equations for the following word equations:

- a. Aluminium + Sulphuric acid \rightarrow Aluminium Sulphate + Hydrogen
- b. Lead Nitrate + Sodium Chloride \rightarrow Sodium Nitrate + Lead Chloride
- c. Calcium + Water \rightarrow Calcium Hydroxide + Hydrogen
- d. Lead Sulphate + Aluminium Hydroxide \rightarrow Aluminium Sulphate + Lead Chloride