



TEST PAPER: CHEMISTRY

Time: 45 Minutes

Class: CBSE 10

Max. Marks: 30 Marks

Date: 18TH July 2018

Marking Scheme: Three questions carry 10 marks each. Questions have 3 subparts each. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.

Question 1:

A. Name the following:

- A metal in liquid state at room temperature
- A non-metal which has lustre
- A metal which has high melting point
- Two allotropes of carbon
- A metal that can be cut with a knife
- A metal which is poor conductor of heat

B. Define the following terms: (any 4)

- | | |
|------------------|----------------|
| a. Metallurgy | d. roasting |
| b. Anodizing | e. calcination |
| c. Galvanization | |

C. Differentiate between acidic and basic oxides.

Question 2:

A. Write equations for the following reactions:

- Iron with steam
- Calcium and potassium with water

B. Give the electron dot structure for the following:

Sodium, magnesium, oxygen, aluminium, chlorine, nitrogen

C. Give reasons:

- Ionic compounds have high melting points.
- Ionic compounds conduct electricity in molten state

Question 2:

A. Show the formation of Na_2O and CaCl_2 with the help of electron dot structure and ionic equations.

B. Show the extraction of sodium by electrolysis with the help of reactions at cathode and anode.

C. What happens when?

- Dilute sulphuric acid is poured on copper plate
- Zinc is added to iron (II) sulphate