



# TEST PAPER: CHEMISTRY

Time: 70 Minutes

Class: CBSE 10th

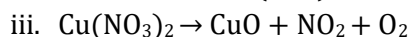
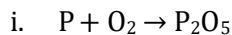
Max. Marks: 50 Marks

Date: 24<sup>th</sup> July, 2018

**Marking Scheme:** Questions carry 10 marks each. Questions have 3 subparts each. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.

## Question: 1

A. Balance the following equations:



B. Write the balanced chemical equations for the following word equations:

i. Calcium hydroxide + carbon dioxide → calcium carbonate + water

ii. Barium chloride + sulphuric acid → barium sulphate + hydrochloric acid

iii. Copper oxide + nitric acid → copper nitrate + water

C. Define the following:

i. Corrosion

ii. Anodizing

## Question: 2

A. Give reasons why an aqueous solution of an acid conduct electricity?

B. Give two important uses of baking soda and washing soda

C. Give reasons for the following:

i. Plaster of Paris should be stored in a moisture-proof container.

ii. Oil and fat containing food items are flushed with nitrogen.

## Question:3

A. Name the following:

i. Medicine used for indigestion

ii. Acid present in ant sting

iii. A metal stored in kerosene oil

B. Show the formation of following ionic compounds by electron transfer

i. NaCl

ii. MgO

D. State four properties of ionic compounds.

## Question: 4

A. Differentiate between exothermic and endothermic reactions. State 3 differences.

B. What is neutralization reaction? Give two examples using balanced reactions.

C. Illustrate the extraction of respective metals from their ores:

i. Zinc from  $ZnCO_3$

ii. Copper from  $Cu_2S$

## Question: 5

A. Differentiate between metals and non-metals. State 3 differences.

B. State two ways of preventing corrosion

C. What are amphoteric oxides? Explain with the help of balanced equations.