



# TEST PAPER: CHEMISTRY

Time: 70 Minutes

Class: CBSE 9<sup>th</sup>

Max. Marks: 50 Marks

Date: 24<sup>th</sup> July, 2018

**Marking Scheme:** Questions carry 10 marks each. Questions have 3 subparts each. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.

## Question: 1

- A. Convert following temperatures into Celsius scale.
- a. 298 K                      b. 410 K                      c. 550 K
- B. What is the physical state of water at:
- a. 298 K                      b. 373 K                      c. 273 K
- C. Give reasons for the following:
- i. Our palm feels cold when we put acetone on it.
- ii. We are able to sip hot tea or milk faster from a saucer rather than a cup.

## Question: 2

- A. Classify following into elements, compounds and mixtures:  
Soil, air, calcium carbonate, calcium, tin, methane, blood and carbon dioxide
- B. Illustrate the method to separate the following
- i. Salt from sea water
- ii. Camphor from salt
- C. Differentiate between homogenous and heterogenous mixture. State 4 differences.

## Question: 3

- A. A solution contains 16 g of urea in 120 g of solution. What is mass by mass percentage of solution?
- B. A solution contains 12 g of sodium chloride in 48 ml of solution. What is mass by volume percentage of solution?
- C. A solution contains 52 g of solute in 208 g of solvent. What is mass by mass percentage of solution?

## Question: 4

- A. Define the following terms:
- i. Colloids
- ii. Tyndal effect
- iii. Polyatomic ion
- B. Give the chemical formulae of following compounds:
- i. Aluminium Chloride
- ii. Potassium Sulphate
- iii. Sodium Hydroxide
- iv. Calcium Carbonate
- C. State the following laws:
- i. Law of conservation of mass
- ii. Law of constant proportion

## Question: 5

- A. Write down the names of following compounds:
- i.  $\text{Al}_2(\text{SO}_4)_3$
- ii.  $\text{KNO}_3$
- iii.  $\text{CaCl}_2$
- B. State the properties of metals and non-metals. Give 3 points each.
- C. i. State the factors affecting evaporation.
- ii. How does evaporation cause cooling?