TEST PAPER: CHEMISTRY

Time: 50 Minutes   **Class:** CBSE 9th

Max. Marks: 30 Marks   **Date:** 26th September, 2018

Marking Scheme: *Questions carry 10 marks each. Questions have 3 subparts each. Subparts (a) and (b) carry 3 marks each and subpart (c) carries 4 marks.*

Question 1:

A. Define the following terms:

   a. Polyatomic ion   b. Atomic mass unit

B. Give the atomicity of following elements: Argon, nitrogen, chlorine, phosphorus, helium, ozone

C. Give the names of elements present in following molecules

   a. Baking powder
   b. Ammonia

Question 2:

A Calculate the molar mass of the following substances. [Na=23, C=12, O=16, N=14, S=32, H=1, Mg=24, Cl=35.5]

   a. Na_2CO_3 d. $Mg(NO_3)_2$
   b. $(NH_4)_2SO_4$ e. CH_3Cl
   c. KNO_3 f. C_2H_5OH

B. What is the mass of following?

   a. 6 moles of water
   b. 2 moles of nitrogen gas

C. Solve the following:

   a. Calculate the number of moles in 128 g of oxygen gas
   b. Calculate the number of moles in 22 g of carbon dioxide

Question 3:

A. State the postulates of Daltons atomic theory.

B. Give the formulae of following:

   a. Aluminium chloride
   b. Magnesium hydroxide
   c. Calcium carbonate

C. Calculate the number of particles of the following:

   a. 46 g of sodium atom [Na=23]
   b. 0.1 mole of carbon atom [C=12]